

N-(3,4-Dichlorophenyl)-2,4-dimethylbenzenesulfonamide

B. Thimme Gowda,^{a*} Sabine Foro,^b P. G. Nirmala^a and Hartmut Fuess^b

^aDepartment of Chemistry, Mangalore University, Mangalagangothri 574 199, Mangalore, India, and ^bInstitute of Materials Science, Darmstadt University of Technology, Petersenstrasse 23, D-64287 Darmstadt, Germany
Correspondence e-mail: gowdabt@yahoo.com

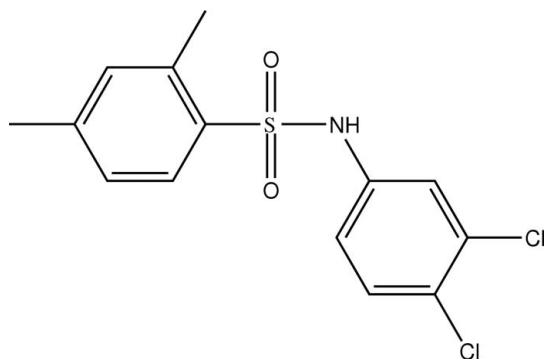
Received 20 July 2009; accepted 21 July 2009

Key indicators: single-crystal X-ray study; $T = 299$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.037; wR factor = 0.100; data-to-parameter ratio = 16.5.

In the crystal structure of the title compound, $\text{C}_{14}\text{H}_{13}\text{Cl}_2\text{NO}_2\text{S}$, the configurations of the N—C bond with respect to the S=O bonds are *trans* and *gauche*. The molecule is bent at the S atom with a C—SO₂—NH—C torsion angle of $-69.7(2)^\circ$. The conformation of the N—H bond is *syn* to the 3-chloro group in the substituted aniline ring. The two benzene rings are tilted with respect to each other by $82.4(1)^\circ$. The presence of N—H...O(S) hydrogen bonding packs the molecules into supramolecular chains along the *b* axis.

Related literature

For our study of the effect of substituents on the structures of *N*-(aryl)-arylsulfonamides, see: Gowda *et al.* (2008; 2009*a,b*). For related structures, see: Gelbrich *et al.* (2007); Perlovich *et al.* (2006).



Experimental

Crystal data

 $\text{C}_{14}\text{H}_{13}\text{Cl}_2\text{NO}_2\text{S}$
 $M_r = 330.21$

Monoclinic, $P2_1/c$
 $a = 8.8046(7)$ Å
 $b = 9.2688(8)$ Å
 $c = 18.947(1)$ Å
 $\beta = 99.644(8)^\circ$
 $V = 1524.4(2)$ Å³

$Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.56$ mm⁻¹
 $T = 299$ K
 $0.44 \times 0.40 \times 0.38$ mm

Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector
 Absorption correction: multi-scan (*CrysAlis RED*; Oxford)

Diffraction, 2009)
 $T_{\min} = 0.790$, $T_{\max} = 0.815$
 10274 measured reflections
 3064 independent reflections
 2618 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.013$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.037$
 $wR(F^2) = 0.100$
 $S = 1.05$
 3064 reflections
 186 parameters
 1 restraint

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.27$ e Å⁻³
 $\Delta\rho_{\min} = -0.39$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1N...O1 ⁱ	0.833 (16)	2.176 (17)	2.984 (2)	164 (2)

Symmetry code: (i) $-x + 2, y - \frac{1}{2}, -z + \frac{1}{2}$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2009); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

BTG thanks the Alexander von Humboldt Foundation, Bonn, Germany, for extensions of his research fellowship.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: TK2510).

References

- Gelbrich, T., Hursthouse, M. B. & Threlfall, T. L. (2007). *Acta Cryst.* **B63**, 621–632.
 Gowda, B. T., Foro, S., Babitha, K. S. & Fuess, H. (2008). *Acta Cryst.* **E64**, o1825.
 Gowda, B. T., Foro, S., Nirmala, P. G., Babitha, K. S. & Fuess, H. (2009*a*). *Acta Cryst.* **E65**, o576.
 Gowda, B. T., Foro, S., Nirmala, P. G. & Fuess, H. (2009*b*). *Acta Cryst.* **E65**, o1940.
 Oxford Diffraction (2009). *CrysAlis CCD* and *CrysAlis RED*. Oxford Diffraction Ltd, Yarnton, England.
 Perlovich, G. L., Tkachev, V. V., Schaper, K.-J. & Raevsky, O. A. (2006). *Acta Cryst.* **E62**, o780–o782.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.

supplementary materials

Acta Cryst. (2009). E65, o1976 [doi:10.1107/S1600536809028840]

N-(3,4-Dichlorophenyl)-2,4-dimethylbenzenesulfonamide

B. T. Gowda, S. Foro, P. G. Nirmala and H. Fuess

Comment

As part of a study of substituent effects on the structures of *N*-(aryl)-arylsulfonamides (Gowda *et al.*, 2008, 2009*a, b*), in the present work, the structure of 2,4-dimethyl-*N*-(3,4-dichlorophenyl)benzenesulfonamide (I) has been determined. The conformations of the N—C bond in the C—SO₂—NH—C segment are *trans* and *gauche* to the S=O bonds (Fig. 1). The molecule is bent at the S atom with the C—SO₂—NH—C torsion angle of -69.7 (2)°, compared to the values of -48.2 (2)° in 2,4-dichloro-*N*-(3,4-dichlorophenyl)benzenesulfonamide (II) (Gowda *et al.*, 2009*b*), and 46.1 (3)° and 47.7 (3)° in the two independent molecules of 2,4-dimethyl-*N*-(phenyl)benzenesulfonamide (III) (Gowda *et al.*, 2009*a*). The conformation of the N—H bond is *syn* to the *meta*-chloro group in the substituted aniline ring. The two benzene rings in (I) are tilted by 82.4 (1)° to each other compared to the values of 68.9 (1)° in II, and 67.5 (1)° and 72.9 (1)° in III. The other bond parameters in (I) are similar to those observed in II, III, and other aryl sulfonamides (Gowda *et al.*, 2008; Perlovich *et al.*, 2006; Gelbrich *et al.*, 2007). The crystal packing of molecules in (I) is *via* N—H⋯O(S) hydrogen bonding (Table 1) leading to a supramolecular chain.

Experimental

A solution of 1,3-xylene (10 ml) in chloroform (40 ml) was treated dropwise with chlorosulfonic acid (25 ml) at 0 °C. After the initial evolution of hydrogen chloride subsided, the reaction mixture was brought to room temperature and poured into crushed ice in a beaker. The chloroform layer was separated, washed with cold water and allowed to evaporate slowly. The residual 2,4-dimethylbenzenesulfonylchloride was treated with 3,4-dichloroaniline in a stoichiometric ratio and boiled for ten minutes. The reaction mixture was then cooled to room temperature and added to ice-cold water (100 ml). The resultant solid, 2,4-dimethyl-*N*-(3,4-dichlorophenyl)benzenesulfonamide, was filtered under suction and washed thoroughly with cold water. It was then recrystallized to constant melting point from dilute ethanol. The prisms used in the X-ray analysis were grown in ethanolic solution by a slow evaporation at room temperature.

Refinement

The H atom of the NH group was located in difference map and was refined with restrained geometry to 0.86 (2) Å. The other H atoms were positioned with idealized geometry using a riding model [C—H = 0.93—0.96 Å], and were refined with isotropic displacement parameters set to 1.2 times of the U_{eq} of the parent atom.

Figures

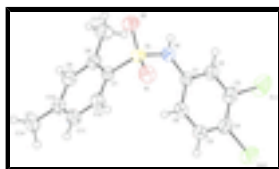


Fig. 1. Molecular structure of (I), showing the atom labelling scheme and displacement ellipsoids are drawn at the 50% probability level

N-(3,4-Dichlorophenyl)-2,4-dimethylbenzenesulfonamide

Crystal data

$C_{14}H_{13}Cl_2NO_2S$	$F_{000} = 680$
$M_r = 330.21$	$D_x = 1.439 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Hall symbol: -P 2ybc	Cell parameters from 4751 reflections
$a = 8.8046 (7) \text{ \AA}$	$\theta = 3.0\text{--}27.6^\circ$
$b = 9.2688 (8) \text{ \AA}$	$\mu = 0.56 \text{ mm}^{-1}$
$c = 18.947 (1) \text{ \AA}$	$T = 299 \text{ K}$
$\beta = 99.644 (8)^\circ$	Prism, colourless
$V = 1524.4 (2) \text{ \AA}^3$	$0.44 \times 0.40 \times 0.38 \text{ mm}$
$Z = 4$	

Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector	3064 independent reflections
Radiation source: fine-focus sealed tube	2618 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.013$
$T = 303 \text{ K}$	$\theta_{\text{max}} = 26.4^\circ$
Rotation method data acquisition using ω and φ scans	$\theta_{\text{min}} = 2.9^\circ$
Absorption correction: multi-scan (CrysAlis RED; Oxford Diffraction, 2009)	$h = -9 \rightarrow 10$
$T_{\text{min}} = 0.790$, $T_{\text{max}} = 0.815$	$k = -11 \rightarrow 11$
10274 measured reflections	$l = -23 \rightarrow 23$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.037$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.100$	$w = 1/[\sigma^2(F_o^2) + (0.0506P)^2 + 0.7353P]$
$S = 1.05$	where $P = (F_o^2 + 2F_c^2)/3$
3064 reflections	$(\Delta/\sigma)_{\text{max}} = 0.007$
186 parameters	$\Delta\rho_{\text{max}} = 0.27 \text{ e \AA}^{-3}$
1 restraint	$\Delta\rho_{\text{min}} = -0.39 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Experimental. CrysAlis RED (Oxford Diffraction, 2009) Empirical absorption correction using spherical harmonics, implemented in SCALE3 ABSPACK scaling algorithm.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.7751 (2)	0.1105 (2)	0.26937 (9)	0.0342 (4)
C2	0.7595 (2)	0.0272 (2)	0.32852 (10)	0.0405 (4)
H2	0.8407	-0.0308	0.3497	0.049*
C3	0.6243 (2)	0.0301 (2)	0.35611 (10)	0.0425 (5)
C4	0.5020 (2)	0.1148 (2)	0.32506 (10)	0.0425 (5)
C5	0.5178 (2)	0.1972 (3)	0.26616 (10)	0.0472 (5)
H5	0.4360	0.2544	0.2449	0.057*
C6	0.6528 (2)	0.1963 (2)	0.23813 (10)	0.0441 (5)
H6	0.6618	0.2529	0.1985	0.053*
C7	0.8600 (2)	0.1276 (2)	0.09846 (9)	0.0334 (4)
C8	0.8733 (2)	-0.0147 (2)	0.07500 (10)	0.0363 (4)
C9	0.7833 (2)	-0.0518 (2)	0.01025 (10)	0.0410 (4)
H9	0.7907	-0.1453	-0.0067	0.049*
C10	0.6828 (2)	0.0432 (2)	-0.03070 (10)	0.0403 (4)
C11	0.6726 (2)	0.1824 (2)	-0.00542 (10)	0.0434 (5)
H11	0.6057	0.2479	-0.0317	0.052*
C12	0.7608 (2)	0.2250 (2)	0.05836 (10)	0.0406 (4)
H12	0.7537	0.3191	0.0746	0.049*
C13	0.9776 (3)	-0.1262 (2)	0.11551 (13)	0.0515 (5)
H13A	0.9396	-0.1519	0.1584	0.062*
H13B	1.0797	-0.0874	0.1277	0.062*
H13C	0.9800	-0.2103	0.0862	0.062*
C14	0.5881 (3)	-0.0036 (3)	-0.10031 (11)	0.0529 (5)
H14A	0.5855	-0.1071	-0.1026	0.064*
H14B	0.6331	0.0336	-0.1394	0.064*
H14C	0.4851	0.0328	-0.1035	0.064*
N1	0.91696 (19)	0.1015 (2)	0.24438 (8)	0.0407 (4)
H1N	0.976 (2)	0.035 (2)	0.2605 (12)	0.049*
O1	0.92732 (17)	0.33947 (15)	0.18616 (7)	0.0471 (4)
O2	1.12802 (16)	0.15602 (19)	0.18124 (8)	0.0530 (4)

supplementary materials

C11	0.61158 (9)	-0.07351 (7)	0.43079 (4)	0.0753 (2)
C12	0.33330 (7)	0.12189 (9)	0.36022 (3)	0.0680 (2)
S1	0.96962 (5)	0.19135 (5)	0.17881 (2)	0.03717 (14)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0378 (10)	0.0348 (9)	0.0301 (8)	-0.0019 (7)	0.0057 (7)	-0.0059 (7)
C2	0.0482 (11)	0.0329 (10)	0.0418 (10)	0.0070 (8)	0.0121 (8)	0.0008 (8)
C3	0.0561 (12)	0.0323 (10)	0.0428 (10)	-0.0013 (9)	0.0190 (9)	-0.0015 (8)
C4	0.0378 (10)	0.0490 (12)	0.0424 (10)	-0.0053 (9)	0.0116 (8)	-0.0117 (9)
C5	0.0372 (11)	0.0635 (14)	0.0389 (10)	0.0079 (9)	0.0010 (8)	-0.0018 (9)
C6	0.0414 (11)	0.0577 (13)	0.0325 (9)	0.0056 (9)	0.0045 (8)	0.0039 (9)
C7	0.0347 (9)	0.0356 (10)	0.0313 (8)	-0.0044 (7)	0.0093 (7)	0.0007 (7)
C8	0.0373 (9)	0.0337 (9)	0.0398 (9)	-0.0012 (8)	0.0121 (7)	0.0019 (7)
C9	0.0476 (11)	0.0346 (10)	0.0421 (10)	-0.0028 (8)	0.0112 (8)	-0.0050 (8)
C10	0.0438 (11)	0.0453 (11)	0.0330 (9)	-0.0057 (9)	0.0100 (8)	-0.0020 (8)
C11	0.0520 (12)	0.0418 (11)	0.0353 (9)	0.0052 (9)	0.0037 (8)	0.0059 (8)
C12	0.0517 (11)	0.0336 (10)	0.0370 (9)	0.0028 (8)	0.0087 (8)	0.0008 (8)
C13	0.0532 (13)	0.0401 (11)	0.0588 (13)	0.0051 (10)	0.0027 (10)	0.0035 (10)
C14	0.0577 (13)	0.0602 (14)	0.0391 (11)	-0.0061 (11)	0.0032 (9)	-0.0050 (10)
N1	0.0381 (9)	0.0497 (10)	0.0350 (8)	0.0079 (7)	0.0080 (7)	0.0061 (7)
O1	0.0542 (9)	0.0374 (8)	0.0481 (8)	-0.0112 (6)	0.0038 (6)	-0.0065 (6)
O2	0.0326 (7)	0.0724 (11)	0.0544 (9)	-0.0062 (7)	0.0087 (6)	-0.0024 (7)
Cl1	0.0973 (5)	0.0599 (4)	0.0818 (5)	0.0169 (3)	0.0531 (4)	0.0292 (3)
Cl2	0.0443 (3)	0.0965 (5)	0.0679 (4)	-0.0024 (3)	0.0234 (3)	-0.0037 (3)
S1	0.0339 (2)	0.0413 (3)	0.0365 (2)	-0.00631 (19)	0.00630 (18)	-0.00184 (19)

Geometric parameters (\AA , $^\circ$)

C1—C2	1.387 (3)	C9—C10	1.389 (3)
C1—C6	1.389 (3)	C9—H9	0.9300
C1—N1	1.410 (2)	C10—C11	1.385 (3)
C2—C3	1.379 (3)	C10—C14	1.501 (3)
C2—H2	0.9300	C11—C12	1.380 (3)
C3—C4	1.382 (3)	C11—H11	0.9300
C3—Cl1	1.729 (2)	C12—H12	0.9300
C4—C5	1.378 (3)	C13—H13A	0.9600
C4—Cl2	1.7283 (19)	C13—H13B	0.9600
C5—C6	1.381 (3)	C13—H13C	0.9600
C5—H5	0.9300	C14—H14A	0.9600
C6—H6	0.9300	C14—H14B	0.9600
C7—C12	1.391 (3)	C14—H14C	0.9600
C7—C8	1.403 (3)	N1—S1	1.6264 (17)
C7—S1	1.7623 (18)	N1—H1N	0.833 (16)
C8—C9	1.387 (3)	O1—S1	1.4353 (15)
C8—C13	1.505 (3)	O2—S1	1.4258 (15)
C2—C1—C6	119.25 (18)	C9—C10—C14	121.03 (19)

C2—C1—N1	116.83 (17)	C12—C11—C10	120.66 (18)
C6—C1—N1	123.92 (17)	C12—C11—H11	119.7
C3—C2—C1	120.29 (18)	C10—C11—H11	119.7
C3—C2—H2	119.9	C11—C12—C7	120.17 (18)
C1—C2—H2	119.9	C11—C12—H12	119.9
C2—C3—C4	120.68 (18)	C7—C12—H12	119.9
C2—C3—C11	118.55 (16)	C8—C13—H13A	109.5
C4—C3—C11	120.77 (15)	C8—C13—H13B	109.5
C5—C4—C3	118.87 (18)	H13A—C13—H13B	109.5
C5—C4—C12	120.05 (16)	C8—C13—H13C	109.5
C3—C4—C12	121.06 (16)	H13A—C13—H13C	109.5
C4—C5—C6	121.20 (19)	H13B—C13—H13C	109.5
C4—C5—H5	119.4	C10—C14—H14A	109.5
C6—C5—H5	119.4	C10—C14—H14B	109.5
C5—C6—C1	119.71 (18)	H14A—C14—H14B	109.5
C5—C6—H6	120.1	C10—C14—H14C	109.5
C1—C6—H6	120.1	H14A—C14—H14C	109.5
C12—C7—C8	121.01 (17)	H14B—C14—H14C	109.5
C12—C7—S1	117.22 (14)	C1—N1—S1	127.44 (14)
C8—C7—S1	121.76 (14)	C1—N1—H1N	117.1 (16)
C9—C8—C7	116.59 (17)	S1—N1—H1N	114.8 (16)
C9—C8—C13	119.36 (18)	O2—S1—O1	119.00 (9)
C7—C8—C13	124.05 (18)	O2—S1—N1	105.08 (9)
C8—C9—C10	123.55 (18)	O1—S1—N1	107.69 (9)
C8—C9—H9	118.2	O2—S1—C7	109.99 (9)
C10—C9—H9	118.2	O1—S1—C7	106.92 (9)
C11—C10—C9	118.01 (18)	N1—S1—C7	107.66 (9)
C11—C10—C14	120.96 (19)		
C6—C1—C2—C3	-0.2 (3)	C8—C9—C10—C11	0.2 (3)
N1—C1—C2—C3	-179.96 (17)	C8—C9—C10—C14	-179.83 (18)
C1—C2—C3—C4	0.5 (3)	C9—C10—C11—C12	0.4 (3)
C1—C2—C3—C11	-179.00 (14)	C14—C10—C11—C12	-179.54 (19)
C2—C3—C4—C5	-0.3 (3)	C10—C11—C12—C7	-0.7 (3)
C11—C3—C4—C5	179.16 (16)	C8—C7—C12—C11	0.3 (3)
C2—C3—C4—C12	-178.74 (15)	S1—C7—C12—C11	179.45 (15)
C11—C3—C4—C12	0.7 (2)	C2—C1—N1—S1	-177.17 (15)
C3—C4—C5—C6	-0.1 (3)	C6—C1—N1—S1	3.1 (3)
C12—C4—C5—C6	178.33 (16)	C1—N1—S1—O2	173.04 (16)
C4—C5—C6—C1	0.4 (3)	C1—N1—S1—O1	45.23 (19)
C2—C1—C6—C5	-0.2 (3)	C1—N1—S1—C7	-69.74 (18)
N1—C1—C6—C5	179.52 (18)	C12—C7—S1—O2	-129.57 (15)
C12—C7—C8—C9	0.3 (3)	C8—C7—S1—O2	49.56 (17)
S1—C7—C8—C9	-178.82 (13)	C12—C7—S1—O1	0.97 (17)
C12—C7—C8—C13	-179.77 (19)	C8—C7—S1—O1	-179.90 (14)
S1—C7—C8—C13	1.1 (3)	C12—C7—S1—N1	116.45 (15)
C7—C8—C9—C10	-0.6 (3)	C8—C7—S1—N1	-64.42 (16)
C13—C8—C9—C10	179.49 (19)		

supplementary materials

Hydrogen-bond geometry (Å, °)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$N1-H1N\cdots O1^i$	0.833 (16)	2.176 (17)	2.984 (2)	164 (2)

Symmetry codes: (i) $-x+2, y-1/2, -z+1/2$.

Fig. 1

